

## Thermochemistry Question Answer Guide

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$Q = m \cdot c \cdot \Delta T$ .  $Q_{\text{glass}} = 1000.0 \cdot 2. (90 - 20) = 14000 \text{ cal}$ .  $Q_{\text{water}} = 1000.1. (90 - 20) = 70000 \text{ cal}$ .  $Q_{\text{calorimeter}} = 70000 + 14000 = 84000 \text{ cal}$ . 1 mol S is 32 g. Molar combustion enthalpy of S is 84000 cal or 84 kcal. Since it is combustion enthalpy;  $\Delta H_{\text{Combustion S}} = -84 \text{ kcal/mol}$ . Thermochemistry Exams and Problem Solutions < Prev.

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Thermochemistry Thermochemistry and Energy and Temperature Thermochemistry is study of changes in energy (heat) associated with physical or chemical changes. Force = push  $F = m \cdot a$  (mass x acceleration) force units: N (newton) =  $\text{kg m s}^{-2}$  Work = force x distance  $W = F \cdot d$  energy units: J (joule) =  $\text{kg m}^2 \text{ s}^{-2}$

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$qH = M \times C \times \Delta T$ . Where M is the mass or volume of the solution being used to test the energy change, C is the Specific Heat Capacity of the solution, ( for example the specific heat capacity of water is  $4.1855 \text{ J g}^{-1} \text{ K}^{-1}$  ), and  $\Delta T$  is the change in temperature.

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## ~~Thermochemistry Question Answer Guide~~

Thermochemistry is a branch of Thermodynamics, a scientific field that comprises the relationships between heat, work, and other forms of energy resulting from different chemical and physical processes. Thermochemistry itself is defined as energy changes when a chemical reaction takes place. Energy is the capacity to supply heat or do work.

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The answer is given per mole of reaction.  $\Delta_f H^\circ (298 \text{ K})$  for  $\text{SF}_4 (\text{g})$ ,  $\text{H}_2\text{O} (\text{l})$ ,  $\text{SO}_2 (\text{g})$  and  $\text{HF} (\text{g})$  are  $-763$ ,  $-286$ ,  $-297$  and  $-273 \text{ kJ mol}^{-1}$ , respectively. What is the value of  $\Delta_r H^\circ (298 \text{ K})$  for the hydrolysis shown above? For this process,  $\Delta_c H^\circ (298 \text{ K}) = -3953 \text{ kJ mol}^{-1}$ .

## ~~Chapter 2: Thermochemistry~~

42. Given that heat neutralization of strong acid and strong base is  $-13.7 \text{ kcal}$ . The heat produced when one mole of HCl is mixed with 0.5 mole of NaOH will be \_\_\_\_\_

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