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especially for salts containing polyatomic ions.

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concentration of acid, as well as the base, are the same. Neutralization Guide Chemistry Answers Neutralization quiz, neutralization quiz questions and answers Page 7/23

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Acid-Base neutralization. One equivalent of acid reacts with one equivalent of base to reach the equivalence point. Normality.

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Number of equivalents per liter of solution.
Normality Formula. Molarity \times (#H⁺ or OH⁻)
Formula when doing a titration with
concentrations in units of Normality:
 $N_a V_a = N_b V_b$.

Chemistry Chapter 18: Neutralization
Reactions Flashcards ...

Solve the neutralization problems below using
the equation $M_a V_a = M_b V_b$. SHOW ALL
WORK: 4) If it takes 54 mL of 0.1 M NaOH to
neutralize 125 mL of an HCl solution, what is
the concentration of the HCl? 5) If it takes
25 mL of 0.05 M HCl to neutralize 345 mL of

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NaOH solution, what is the concentration of the NaOH solution?

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The full ionic equation for the
neutralization of hydrochloric acid by sodium
hydroxide is written as follows: (4.4.2) $H +$

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$(a q) + Cl^- (a q) + Na^+ (a q) + OH^- (a q) \rightarrow Na^+ (a q) + Cl^- (a q) + H_2O (l)$ Since the acid and base are both strong, they are fully ionized and so are written as ions, as is the NaCl formed as a product.

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