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Course Chemistry #10

~~Introduction to Oxidation~~

~~Reduction (Redox) Reactions~~

Oxidation and Reduction

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Step Example Oxidation and

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~~[Workshop Recording]~~ **Rules**

for Oxidation Numbers

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Chemistry - Oxidation and

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#32 (Higher Tier) Redox

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method Birch Reduction
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Cell Potential \u0026

Notation, Redox Half

Reactions, Nernst Equation

8 OXIDATION N REACTION 1

Oxidation and Reduction

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~~Question of the Day: Answers~~

~~Oxidation, Reduction, and
Assigning Oxidation Numbers~~

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In oxidation/reduction reactions, there is a transfer of charge - an

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Electrochemistry Answers
Electron - from one species
to another. Oxidation is the
loss of electrons and
reduction is a gain in
electrons. Use these
acronyms to help you
remember: Leo Ger - L oss of
e lectrons o xidation; G ain

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Oxidation Reduction And
of electrons (reduction or
Oil Rig - Oxidation i
nvolves loss, Reduction i
nvolves gain - of
electrons.)

**Review: Oxidation/Reduction
(Redox) Reactions and ...**

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Oxidation & Reduction In
Terms Of Oxidation Number
(O.N) According to this
concept, the increase in
oxidation number means
Oxidation (loss of the
electron) whereas the
decrease in oxidation number

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means Reduction (gain of
electron). $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$

The O.N of Zn increased from
0 to 2 i.e Zn got oxidized.

**The Concept Of Oxidation &
Reduction | Notes, Videos,**

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Oxidation and reduction can also be defined in terms of changes in composition. The original meaning of oxidation was “adding oxygen,” so when oxygen is added to a molecule, the

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molecule is being
oxidized. The reverse is true
for reduction: if a molecule
loses oxygen atoms, the
molecule is being reduced.
For example, the
acetaldehyde
(CH_3CHO) molecule

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takes on an oxygen atom to
...

**5.6: Oxidation-Reduction
(Redox) Reactions -
Chemistry ...**

The loss of electrons is called oxidation. The gain

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of electrons is called
reduction. Because any loss
of electrons by one
substance must be
accompanied by a gain in
electrons by something else,
oxidation and reduction
always occur together. As

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such, electron-transfer reactions are also called oxidation-reduction reactions, or simply redox reactions.

6.1: Oxidation-Reduction (Redox) Reactions -

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Answers maintain our
civilization and the
metabolism of foods that
furnish the energy that
keeps us alive both involve

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also redox reactions. 05 CTR
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517 THE

**Oxidation Reduction Concept
Review Answers**

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classical concept of
oxidation and reduction

According to classical
concept following

definitions were proposed to
explain the process of
oxidation and reduction.

Oxidation: It is a process

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of chemical addition of
oxygen or any
electronegative radical or
removal of hydrogen or any
electropositive radical.

**Classical Concept of
Oxidation and Reduction -**

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Oxidation-reduction reactions are of central importance in organic chemistry and biochemistry. The burning of fuels that provides the energy to maintain our civilization

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and the metabolism of foods
that furnish the energy that
keeps us alive both involve
redox reactions. The carbon
compounds in organic and
biochemistry tend to be
covalent molecules.

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5.6: Redox Reactions in Organic Chemistry and Biochemistry

This page looks at the various definitions of oxidation and reduction (redox) in terms of the transfer of oxygen, hydrogen

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and electrons. It also
explains the terms oxidising
agent and reducing agent. An
oxidising agent is substance
which oxidises something
else. In the above example,
the iron(III ...

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**DEFINITIONS OF OXIDATION AND
REDUCTION (REDOX)**

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reviewing a book oxidation
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Oxidation And Reduction

Oxidation : Oxidation is a process in which an atom taking part in chemical reaction loses one or more electrons. The loss of

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Electrochemistry Answers provides the energy to maintain our civilization and the metabolism of foods that furnish the energy that keeps us alive both involve redox reactions.

5.7: Redox Reactions in

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Electrochemistry ...

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"removal of oxygen".

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the gain of O or loss of H.

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Reduction is the loss of O or gain of H . Oxidation and reduction always occur together, even though they can be written as separate chemical equations. Concept Review Exercises

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