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4.1 DEFINING THE ATOM (pages
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theories of matter and
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the tiny size of individual

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The sum of the protons and
neutrons in the nucleus of
an atom. $MASS\# = \#$
protons (atomic number) +

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$\# \text{neutrons} = \text{Mass} \# - \# \text{Protons (atomic number)}$
 $\# \text{Protons} = \text{Mass} \# - \# \text{Neutrons}$. ISOTOPES: Are atoms of the same element that have different numbers of neutrons and different mass numbers.

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mass. atomic mass unit (amu)

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atomic number (Z) the
smallest unit of an element
that maintains the
properties... the weighted
average of the masses of the
isotopes of an eleme... a
unit of mass equal to one-
twelfth the mass of a

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Answers

2) All atoms of a given element are identical in size, mass, and chemical properties. 3) Atoms of a specific element different

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from those of another
element. 4) Different atoms
combine in simple whole-
number ratios to form
compounds. 5) A chemical
reaction, atoms are
separated combined or
rearranged.

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